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4 *P67130A0428* 3 Ammonium nitrate is very

soluble in water. $\text{NH}_4\text{NO}_3(\text{s}) + \text{aq} \rightarrow \text{NH}_4^+(\text{aq}) + \text{NO}_3^-(\text{aq})$ $\Delta H_{-1} = +25.8 \text{ kJ mol}^{-1}$ What is the best explanation ...

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(i) Considering that the flask gets cooler as the reaction proceeds, what drives the chemical reaction between $\text{NaHCO}_3(\text{s})$ and $\text{HC}_2\text{H}_3\text{O}_2(\text{aq})$? Answer by ...